{youtube}xirPkCI1sMA{/youtube}

How to find the average atomic mass of an element. You need to know the mass of each isotope and the percent (%) abundance of each as well. Multiply each mass by its corresponding percentage, and add these products together.

Answer:

Each percentage here is called the **isotopic abundance** of that particular isotope.

The **average atomic mass** (mass on the periodic table) is a weighted average of all isotopes.

Average Atomic Mass = (%□ *of isotope* 1)(*mass of isotope* 1)
+ (%□ *of isotope* 2)(*mass of isotope* 2) + ...

In this case, the average atomic mass of magnesium is

(0.7870)(23.985 amu) + (0.1013)(24.985 amu) + (0.1117)(25.983 amu)
= **24.309 amu**

□□□□□

http://www.youtube.com/

http://lessons.chemistnate.com/