

Gibbs energy change and equilibrium

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{youtube}7YoNyTWMNkQ?{/youtube}

In this video explains In thermodynamics, the Gibbs free energy IUPAC recommended name: Gibbs energy or Gibbs function; also known as free enthalpy to distinguish it from Helmholtz free energy is a thermodynamic potential that measures the "usefulness" or process-initiating work obtainable from a thermodynamic system at a constant temperature and pressure isothermal, isobaric. Just as in mechanics, where potential energy is defined as capacity to do work, similarly different potentials have different meanings. The Gibbs free energy SI units kJ/mol is the maximum amount of non-expansion work that can be extracted from a thermodynamically closed system one that can exchange heat and work with its surroundings, but not matter; this maximum can be attained only in a completely reversible process.

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